**In class, no-calculations mini-exam 2: Chem 6854 Spring 2016**

Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_/36 pts

 2 pts/problem

1. **If φ = Ceikx, express <φ\*|φ> as an integral over all space from 0🡪L. (Don’t evaluate: just set up)**
2. **If we were to `normalize φ’ above, what must the expression above be set equal to ??**
3. **Briefly explain why spherical (r, θ,φ) rather than Cartesian (x,y,z) coordinates are used to solve the**

**Hydrogen atom Schrödinger equation.**

1. **The energy of an H atom is proportional to:**
2. -n2 b) -L(L+1)/n2 c) –mL2(L)(L+1)/n2 d) -L2/n2 e) -1/n2
3. **If = kinetic energy operator and +½ kx2 = the potential operator for a harmonic oscillator system, E = the eigen energy and ψ the wave function write down the basic form of the Schrödinger eigenvalue problem we need to solve.**
4. **The quantum energy of a rigid rotor varies with:**
5. J(J+1) b) 1/J c) J d) J2/(1+J) e) J-1
6. **The correct expression for the energy of a quantum harmonic oscillator is: (**ω **=f/2**π **radians)**
7. En = hωn b)En = hω(n +1) c) En = h(n + ½ )/ω d) En = hω(n + ½) e) En = μr2n(n+1)
8. **Given that n=2 for an H atom, what are possible values of L ?**
9. 2,1,0 b) 2,1,0,-1,2 c) 1,0 d)1,0,-1 e) 2
10. **If L= 1, what are the possible values of mL ?**
11. 1,0 b) 1,0,-1 c) 2,1,0 d) 0,-1 e) 0
12. **If L=3, what corresponding orbital equivalent are we talking about ?**
13. s orbits b) p orbits c) d orbits d) f orbits
14. **Which term below represents the main reason an exact quantum solution to multi-electron systems of either atoms or molecules is essentially impossible:**
15. **electron-nuclear interaction b) nuclear-nuclear interaction c) electron-electron interaction d) kinetic E**
16. **In the variational method of approximation used in the Hartree-Fock calculations you carried out with the Spartan QChem program, what critical basis set is used throughout:**
17. **Roothan-Hall wavefunctions b) Slater orbitals (STO) c) H1s orbitals d)Laguerre polynomials**
18. **Given that an approximated wave function ψapprox =ψ has the form below, what does the variational method**

**minimize? ψapprox = c1φ1(r1,θ1 ) + c2φ2(r2,θ2)**

1. d<ψ|ψ>/dE =0 c) dφ1/dc1= dφ2/dc2 =0
2. b) d(<ψ|ψ>/<ψ|ψ>)/dc1= d(<ψ|ψ>/<ψ|ψ>)/dc2= 0 d) d) dψ/dr1 = dψ/dr2 = 0

**14) The secular equation for a two basis element ( φ1 and φ2) approximation of Ψ for a given quantum Hamiltonian is written below :**

det

= 0

**a) define H11= H11 using bra-ket notation (1 pt)**

**b) define S12= S12 using standard integral notation. Assume the integration range is 0🡪∞. (1 pt)**

**c) The secular equation above arises from solving the matrix problem below :**

0

0

C1

C2

=

**Why is the determinant (det) of the matrix set =0 ? (2 pts)**

1. **Briefly describe the `correlation’ problem inherent in the Hartree-Fock methodology used in our Spartan QChem program**
2. **Recalling Professor Simpson’s two lectures, what prime advantage does the DFT approach to quantum calculation have over other post Hartree-Fock methods ?**
3. The Universal function is known

b) Computation time goes up only as N3, N=# basis elements (cheaper)

 c) DFT leaves out electron-electron interactions

 d) The Born-Oppenheimer approximation is not applied

1. **How many angels can dance on the head of a pin ?**
2. get bent, Fong b) cats rule, you drool c) 4n+2 d) Nookie !