**Mini-quiz #14 Chemistry 1114 Monday 8 October 2012**

**6 pts**

Your name:\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Assuming the following atomic masses/mol:

C=12 H=1 O=16 Cl=35.5 P=31 Cu=63.5 Ca=40

Compute the molecular masses/mol for the compounds below:

14.1 CuCO3 \_\_\_\_\_\_\_\_123.5\_\_\_\_\_\_\_\_\_\_\_\_g/mol

**Cu + C + 3\*O**

63.5 + 12 +3\*16=123.5

14.2. PCl5 \_\_\_\_\_\_208.5\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ g/mol

**P +5\*Cl**

31+5\*35.5=208.5

14.3. CaCl2 \*3H2O \_\_\_\_\_\_\_165\_\_\_\_\_\_\_\_\_ g/mol

**Ca + 2\*Cl +2\*(2H + O)**

40 + 2\*35.5 + 3\*(2\*1+16) = 165