**CHEM 1114**

**Challenge problem 2**

**Due Wednesday 19 September 2012 in class**

**(4 pts)**

**2. The energy of a photon = ΔE=hf =hc/λ in Joules**

**From your text (p. 72) the electronic transitions in an H atom obey Bohr’s derived equation:**

**ΔE = -2.187\*10-18nfinal2 – 1/ninitial2)**

**What is the wavelength of light (in meters)**

**Associated with ninitial = 3🡪nfinal = 1 ?**