**H X**

**| |**KOH or KOR /CCl4 or ROH (solvent)

***DEHYDROHALOGENATION:*** C-C --------------> C=C

E2-the facts β α

**action or effect E11 E2 (main route)**

**[substrate] increased**  Rate increases

**[OH-] increased**  Rate Increases

**Rearrangement ?**  NO

**Primary H effects ?**  YES

**halogen (`element’) effect**  RI >>RBr>>RCl>>RF

**on rate ~(25,000:500: 10:1) relative rates**

**effect of order\* on rate**  3o >2o >1o

(\*order of α carbon)

**product alkene distribution** follows Zaitsev Rule

*other `facts’ on dehydrohalogenation of an alkyl halide:*

0) βC must have at least 1 H for E2; E1 happens even without β-H

1) OH- is consumed

2) E2 is generally the favored route. E1 seen

***The Mechanisms:***(new C—C π bond forms)

**| | | | | |**

E2: OH- + -C-C- --> -C-C- -------------> -C=C- + H2O + X-

*one step* | | (C-X breaks)

H XH[+]X(-)

OH[-] *both leave ~ simultaneously*

(C-H breaks)

*transition state*

β position

**supplement #14: *ELMINATION PATHWAYS***

***CHEMISTRY 3514 Alfred State College***

***DEHYDRATION:*  H OH**

**| | *H+*** (acid)

C-C --------------> C=C

**E1-**The facts β α

**action or effect E1 (main route)**

**[substrate] increased** Rate increases

**[H+] increased**  Rate increases

**Rearrangement ?** YES

**Primary H effects ?** NO

**effect of order\* on rate** 3o >2o >1o

(\*order of α carbon)(sometimes 1o goes E2)

**product alkene distribution**  follows Zaitsev Rule f

***other `facts’ on dehydration***

0) β C doesn’t need H for E1 to run

1) H+ is NOT consumed

2) E1 is ALMOST always the favored route

3) carbocation can also undergo substitution (a competitive situation)

*acid catalyzed E1*

**| | rev. | |**

H+ + -C-C- -C-C- protonation [equilibrium step=>

**| | | | reversible (rev.)**

**β position**H OH **H OH2 +**

**| | | | | | fast | |**

***E1*** -C-C- --> -C-C- + **H2O****-**C-C------> -C=C- + H+

**| | | + + |**

H OH2+ **H H**

*carbocation* *rearranged carbocation*

***step 1 step 2***

note: when the β Carbon has no H...step 2 comes after rearrangement to a carbocation with at least one neighboring

Methyl. This explains why dehydration occurs on 2,2,4,4-tetramethyl-3-pentanol to make 2,3,4,4-tetramethyl2-pentene as its chief product.