**Homework #9 Chemistry 1114 section 2 (Fong) due Fri 6 Oct 2017 8 pts (in class) Show your work !!**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**1) 2.0 grams of C are burned to make 7.333 g COx. What is the empiric formula for COx ? (at. mass C=12 g/mol; at mass O=16 g/mol)**

**Oxygen mass=7.333-2=5.333 g=> 5.333/16=mol O=0.333; mol C= 2.0 g/12=0.1666**

**Mol O/mol C= 0.333/0.166=2=> CO2**

**\_\_CO2\_\_\_\_ COx formula**

**2) CaHb is burned in O2 to make 0.1111 g CO2 and 0.02273 g H2O.**

**What is the empiric formula for CaHb ? (mol. mass CO2 =44 g/mol;**

**mol. mass H2O=18 g/mol)**

**0.111 g CO2/44=0.00252 mol ; mol C/mol CO2 =1/1=x/0.00252=> x=0.00252**

**0.02273 g H2O/18=0.00126 mol ; mol H/mol H2O=2/1= y/0.00126=>y=2\*0.00126**

**= 0.00252**

**Mol O/mol H=x/y= 0.00252/0.00252=1**

**\_\_\_\_CH\_\_ CaHb formula**

**3) Balance me:**

**\_2\_C8H18 + \_25\_\_O2🡪 \_16\_\_CO2 + \_18\_\_H2O**

**4) A 4 gram sample of an iron oxide is heated in vacuum to produce 3.109 g of pure**

**iron (Fe). What is the original oxide’s formula ? (at. wt of Fe=55.845 g/mol; at.**

**wt of O=16.0 g/mol)**

**FeOx🡪 Fe + O**

**Sample mass 4.00 g 3.109 g 4.000-3.109=0.891 g**

**At. wt 55.845 16.00**

**Mol 3.109/55.845=0.0557 0.891/16=0.0557**

**Fe mol/O mol = 0.0557/0.0557=1=> FeO**

**\_\_FeO\_\_\_\_ Fe oxide formula**