**Homework #10 Chemistry 1114 due Wednesday 11 Oct 2017 (in class) Show your work !!**

**Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_/6**

**MW(g/mol): 44 32 44 18**

**Given the balanced equation: C3H8 + 5O2 🡪 3CO2 + 4H2O**

**1) How many grams of O2 are burned if 0.0375 mol of CO2 are produced?**

**\_\_\_\_\_ g O2**

**2) How many grams of C3H8 are burned to produce 9 g CO2 ?**

**\_\_\_\_\_\_ g C3H8**

**3) How many molecules of CO2 are produced when we burn 2.444 g C3H8 ?**

**\_\_\_\_\_\_\_\_\_\_\_ molecules CO2**