**Mole HomeWork 8**

**Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_/5**

**MW(g/mol): 330 32 44 18**

**C11H22O11 + 11O2🡪 11CO2 +11H2O**

**1) Your body metabolizes sucrose (sugar=C11H22O11)) according to the balanced**

**reaction written above.**

**You drink a 12 oz can of Coke which contains 44 g of sucrose. Your `yield’ of CO2**

**After metabolizing this amounts to 58.08 g. What is the efficiency (% yield) your**

**body’s metabolism is operating at ?**

**\_\_\_\_\_\_\_\_ % (2 pts)**

**2) Write the balanced and complete molecular, complete ionic and net ionic**

**equation for the reaction between Na3PO4(aq) and AgNO3(aq) . Note that Ag3PO4**

**forms a precipitate (e.g. is written as Ag3PO4(s)) . All other salts are aqueous**

**(3 pts)**

**complete molecular**

**complete ionic**

**net ionic**