**Mole HomeWork 5**

**Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Do these problems and turn your completed work to class**

**Friday 9 October. You can put your work on a separate piece of paper or copy this off and do the work in the spaces provided. SHOW WORK or NO CREDIT**

**1) 2.0 grams of C are burned to make 7.333 g COx. What is the empiric formula for COx ? (at. mass C=12 g/mol; at mass O=16 g/mol)**

**\_\_\_\_\_\_\_\_\_ COx formula**

**2) CaHb is burned in O2 to make 0.1111 g CO2 and 0.02273 g H2O.**

**What is the empiric formula for CaHb ? (mol. mass CO2 =44 g/mol;**

**mol. mass H2O=18 g/mol)**

**\_\_\_\_\_\_\_\_\_ CaHb formula**

**3) Balance me:**

**\_\_C8H18 + \_\_\_O2🡪 \_\_\_CO2 + \_\_\_H2O**

**4) How many grams of O2 are consumed to make 8.25 g CO2**

**In the reaction: C3H8 + 5O2 🡪 3CO2 + 4 H2O**

**mol. Mass C3H8=44; mol mass O2=32; mol. mass CO2 =44 g/mol; mol. mass H2O=18 g/mol)**

**\_\_\_\_ g O2**