**HomeWork 9**

**Due Wednesday 10/21/15**

**Your name: \_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_/7**

**1) In the Arrhenius model, a base is a(n) \_\_\_\_\_\_OH- donor\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**2) In the Bronsted model. A base is a(n) \_\_\_\_\_\_\_proton acceptor\_\_\_\_\_\_\_\_\_\_\_\_\_**

**3) Write the hydrolysis reaction for HSiO3-:**

**HSiO3- + H-OH🡪 H2SiO3 + OH-**

**4) Identify the acid (A), base (B) , conjugate acid (CA) and conjugate base (CB)**

**in the reaction below:**

**(2 pts)**

**H2PO3- + H2PO42- 🡪 H3PO3 + HPO43-**

**I am: \_B\_\_\_ \_A\_\_ \_\_CA\_ \_CB\_\_**

**5) In the Arrhenius model: Acid + Base🡪 \_\_salt\_\_\_\_\_\_\_ + \_\_water\_\_\_\_\_\_\_\_\_\_\_**

**6) In the Bronsted model: Acid + Base🡪 \_conjugate base\_\_\_ + \_\_\_conjugate acid\_\_**