**Homework #6 Chemistry 1114 section 2 (Fong) due Mon 5 March 2018 12 pts (in class)**

**Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**The molecular mass of crystal meth (C10H13N) is 149 g/mol.**

**Assuming the atomic weights of C=12 g/mol; H=1 g/mol and N=14 g/mol**

**and 1 mole count=6\*1023. SHOW WORK OR NO CREDIT (3 pts per problem)**

**6.1. How many grams of crystal meth are formed from 0.4362 g H?**

**\_\_\_\_\_\_\_\_ g meth**

**6.2 How many grams of C are found in 1.6666 mol crystal meth?**

**\_\_\_\_\_\_\_ g C**

**6.3. How many atoms of N are present if the meth sample contains 20 g of C ?**

**\_\_\_\_\_\_\_\_\_\_\_ atoms of N**

**6.4. How many grams of C are in a sample of meth which contains 1.3\*1023 atoms of H ?**

**\_\_\_\_\_\_\_\_\_\_\_ grams C**