**Mini-quiz #23 Chemistry 1114 section 2 (Fong) 24 October 2014 4 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

If an element is oxidized it \_\_\_\_\_\_loses\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons

2) What is the oxidation number of each Cu in Cu2O ? \_\_+1\_\_\_\_\_

3) What is being oxidized and reduced in the reaction below ?

**MgO + H2(g) 🡪 Mg + H2O**

**Oxidized = \_\_\_\_\_H2\_\_\_\_\_**

**Reduced= \_\_\_\_\_Mg in MgO\_\_\_\_**

**Mini-quiz #23 Chemistry 1114 section 2 (Fong) 24 October 2014 4 pts B**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

1) What is the oxidation number of Ni in in NiO3 ? \_\_\_\_+6\_\_\_

1. If an element is reduced it \_\_\_\_gains\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons

3) What is being oxidized and reduced in the reaction below ?

**Cu2O + Mg 🡪 MgO + 2Cu**

**Oxidized = \_\_\_Mg\_\_\_\_\_\_\_**

**Reduced= \_\_\_\_Cu in Cu2O\_\_\_\_\_**

**Mini-quiz #23 Chemistry 1114 section 2 (Fong) 24 October 2014 4 pts C**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

1) What is being oxidized and reduced in the reaction below ?

**Ag2O + Zn 🡪 ZnO + 2Ag**

**Oxidized = \_\_\_\_Zn\_\_\_\_\_\_**

**Reduced= \_\_\_\_Ag in Ag2O\_\_\_\_\_**

2) What is the oxidation number of Mn in in MnO3 ? \_\_\_6\_\_\_\_

1. If an element is reduced it \_\_\_\_\_\_\_gains\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons