**Mini-quiz #20 Chemistry 1114 section 2 (Fong) 17 October 2014 3 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Gasoline, C8H18, burns according to the balanced equation:**

 **2C8H18 + 25 O2 🡪 16CO2  + 18H2O**

**MW 114 g/mol**

**A 22.8 g sample of gasoline is experimentally burned in excess O2 to produce 0.8 moles of CO2.**

**What is the % yield for the reaction ?**

 **` \_\_\_\_\_\_\_\_\_\_\_% yield**

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**Gasoline, C8H18, burns according to the balanced equation:**

 **2C8H18 + 25 O2 🡪 16CO2  + 18H2O**

**MW 114 g/mol**

**A 11.4 g sample of gasoline is experimentally burned in excess O2 to produce 0.3 moles of H2O.**

**What is the % yield for the reaction ?**

 **` \_\_\_\_\_\_\_\_\_\_\_% yield**

**Mini-quiz #20 Chemistry 1114 section 2 (Fong) 17 October 2014 3 pts C**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Pentane, C5H112, burns according to the balanced equation:**

 **C5H12 + 8 O2  🡪 5CO2  + 6H2O**

**MW 72 g/mol**

**An 18 g sample of pentane is experimentally burned in excess O2 to produce 0.375 mol H2O.**

**What is the % yield for the reaction ?**

 **` \_\_\_\_\_\_\_\_\_\_\_% yield**