**Mini-quiz #19 Chemistry 1114 section 2 (Fong) 15 October 2014 3 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Propane (C3H8) burns according to the stoichiometrically balanced reaction below:**

**C3H8 +5 O2 🡪 3CO2 + 4H2O**

**MW (g/mol) 44 32 44 18**

**How many moles of CO2  can form if we burn 22 grams of C3H8 and 32 g of O2 ? (Show work !)**

**moles CO2 =\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Mini-quiz #19 Chemistry 1114 section 2 (Fong) 13 October 2014 4 pts B**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Propane (C3H8) burns according to the stoichiometrically balanced reaction below:**

**C3H8 +5 O2 🡪 3CO2 + 4H2O**

**MW (g/mol) 44 32 44 18**

**a) How many moles of H2O can form if we burn 88 grams of C3H8 and 80 g O2? (Show work !)**

**Moles H2O =\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Mini-quiz #19 Chemistry 1114 section 2 (Fong) 13 October 2014 4 pts C**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Propane (C3H8) burns according to the stoichiometrically balanced reaction below:**

**C3H8 +5 O2 🡪 3CO2 + 4H2O**

**MW (g/mol) 44 32 44 18**

**How many moles of CO2 can form if we combine 88 grams of C3H8 and 64 grams of O2?**

**Moles CO2 \_\_\_\_\_\_\_**