**Mini-quiz #18 Chemistry 1114 section 2 (Fong) 13 October 2014 4 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Propane (C3H8) burns according to the stoichiometrically balanced reaction below:**

**C3H8 +5 O2 🡪 3CO2 + 4H2O**

**MW (g/mol) 44 32 44 18**

**a) How many grams of CO2 form when we burn 0.666 grams of C3H8 ? (Show work !)**

**g CO2 =\_\_\_\_\_\_\_\_\_\_\_\_\_**

**b)How many moles of H2O form with 4.5\*1023 molecules of CO2. (1 mol count = 6\*1023 molecules)**

**Mol H2O=\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Mini-quiz #18 Chemistry 1114 section 2 (Fong) 13 October 2014 4 pts B**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Propane (C3H8) burns according to the stoichiometrically balanced reaction below:**

**C3H8 +5 O2 🡪 3CO2 + 4H2O**

**MW (g/mol) 44 32 44 18**

**a) How many grams of H2O form when we burn 2.444 grams of C3H8 ? (Show work !)**

**g CO2 =\_\_\_\_\_\_\_\_\_\_\_\_\_**

**b)How many moles of CO2 form from 2\*1024 molecules ofC3H8. (1 mol count = 6\*1023 molecules)**

**Moles CO2=\_\_\_\_\_\_\_\_\_\_\_\_\_**