**Mini-quiz #17 Chemistry 1114 section 2 (Fong) 10 October 2014 3 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Butane (C4H10) has a molecular weight of 58 g/mol and burns according to the stoichiometrically balanced reaction below:**

**2C4H10 +13 O2 🡪 8CO2 + 10H2O**

**How many moles of CO2 form when we burn 29 grams of C4H10 ? (Show work !)**

**Mol CO2 =\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Mini-quiz #17 Chemistry 1114 section 2 (Fong) 10 October 2014 3 pts B**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Pentane (C5H12) has a molecular weight of 72 g/mol and burns according to the stoichiometrically balanced reaction below:**

**C5H12 +8 O2 🡪 5CO2 + 6H2O**

**How many moles of H2O form when we burn 12 grams of C5H12 ? (Show work !)**

**Moles H2O=\_\_\_\_\_\_\_\_\_\_**

**Mini-quiz #17 Chemistry 1114 section 2 (Fong) 10 October 2014 3 pts C**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Hexane (C6H14) has a molecular weight of 86 g/mol and burns according to the stoichiometrically balanced reaction below:**

**2C6H14 +19 O2 🡪 12CO2 + 14H2O**

**How many moles of H2O form when we burn 36.855 grams of C6H14 ? (Show work !)**

**Moles H2O=\_\_\_\_\_\_\_\_\_\_\_\_\_**