**Mini-quiz #15 Chemistry 1114 section 2 (Fong) 3 October 2014 4 pts A**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**SHOW WORK FOR COMBUSTION PROBLEM OR NO CREDIT WILL BE ASSIGNED**

A hydrocarbon with the formula CxHy is burned to form 11 grams CO2 and 4.5 grams of H2O.

Given the molecular weights : CO2 = 44 g/mol, H2O=18 g/mol, provide a whole-numbered formula for CxHy. (2 pts)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ empiric formula for CxHy

**Balance us:**

**\_\_\_\_H2O2 + \_\_\_\_Cu -🡪 \_\_\_Cu(OH)2 +\_\_\_H2O + \_\_O2**

**\_\_\_\_C6H14 + \_\_\_\_\_O2🡪 \_\_\_\_\_CO2 + \_\_\_\_H2O**

**Mini-quiz #15 Chemistry 1114 section 2 (Fong) 3 October 2014 4 pts B**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**SHOW WORK FOR COMBUSTION PROBLEM OR NO CREDIT WILL BE ASSIGNED**

A hydrocarbon with the formula CxHy is burned to form 5.5 grams CO2 and 1.125 grams of H2O.

Given the molecular weights : CO2 = 44 g/mol, H2O=18 g/mol, provide a whole-numbered formula for CxHy. (2 pts)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ empiric formula for CxHy

**Balance us:**

**H2O2 + Cu -🡪 \_\_\_Cu(OH)2 +\_\_\_H2O + \_\_O2**

**\_\_\_\_C4H10 + \_\_\_\_\_O2🡪 \_\_\_\_\_CO2 + \_\_\_\_H2O**