**Mini-quiz #20 Chemistry 1114 Wednesday 6 November 2013**

**4 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ A**

**Using your Periodic Table (or one provided):**

1. Draw the best Lewis structure for SiOCl2 assuming that both Cl and O are attached to the central Si, and, that the octet rule cannot be broken.
2. Draw the best Lewis structure for PO43- assuming that all the O are attached to a central P, and, that you can break the octet rule to minimize formal charge.

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**4 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ B**

**Using your Periodic Table (or one provided):**

1. Draw the best Lewis structure for COF2 assuming that both F and O are attached to the central Si, and, that the octet rule cannot be broken.
2. Draw the best Lewis structure for SO42- assuming that all the O are attached to a central S, and, that you can break the octet rule to minimize formal charge.