**Mini-quiz #18 Chemistry 1114 Friday 1 November 2013**

**6 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ A**

Given the balanced reaction below:

(MW g/mol)= 114 32 44 18

2C8H18 + 25 O2 🡪 16 CO2 + 18 H2O

1. What is the limiting reagent to make H2O if you burn 7.03 g C8H18 with

1.388\*1024 molecules of O2 ? [1 mol count=6\*1023 molecules].

Circle your choice

Limiting Reagent = **O2 C8H18**  3 pts

1. How many moles of CO2 can you make if you burn 0.5 grams O2 with 0.3 grams C8H18 ??

Mol CO2 = \_\_\_\_\_\_\_\_\_\_\_\_\_3 pts

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**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ B**

(MW g/mol)= 114 32 44 18

2C8H18 + 25 O2 🡪 16 CO2 + 18 H2O

1. What is the limiting reagent to make H2O if you burn 28.12 g C8H18 with

1.388\*1024 molecules of O2 ? [1 mol count=6\*1023 molecules].

Circle your choice

Limiting Reagent = **O2 C8H18** 3 pts

1. How many moles of CO2 can you make if you burn 1.0 grams O2 with 0.6 grams C8H18 ??

Mol CO2 = \_\_\_\_\_\_\_\_\_\_\_\_\_3 pts