**Mini-quiz #16 Chemistry 1114 Friday 18 October 2013**

 **7 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ A**

A sample of CxHy is burned in pure O2 to form 13 grams of CO2 and 6.6477 grams of H2O. Given that the molecular weight of CO2 =44 g/mol and the molecular weight of H2O=18 g/mol, determine the empiric formula for the CxHy sample (3 pts)

 CxHy =

Balance me ! (1 pt for each correct coefficient):

 \_\_\_\_C4H10 + \_\_\_\_ O2 🡪 \_\_\_CO2 + \_\_\_H2O

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 **7 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ B**

A sample of CxHy is burned in pure O2 to form 6.500 grams of CO2 and 3.989 grams of H2O. Given that the molecular weight of CO2 =44 g/mol and the molecular weight of H2O=18 g/mol, determine the empiric formula for the CxHy sample (3 pts)

CxHy =

Balance me ! (1 pt for each correct coefficient):

 \_\_\_\_C5H12 + \_\_\_\_ O2 🡪 \_\_\_CO2 + \_\_\_H2O