**Mini-quiz #15 Chemistry 1114 Wednesday 16 October 2013**

**6 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ A**

1. Circle the compound formulas below that are not empiric formulas:

**CHO C2H4O4 N21F7O3 H4S2O16**

1. Compound **A** 21.82 % C, 72.72% O and 5.45 % H by weight. Given that the atomic masses of C=12 g/mol, O=16 g/mol and H= 1 g/mol, determine the empiric formula for compound **A**.

EMPIRIC FORMULA OF COMPOUND A =\_\_\_\_\_\_\_\_\_\_\_\_

1. The molecular weight of compound **A** above is 550 g/mol. What is the molecular formula for compound **A** ?

MOLECULAR FORMULA OF COMPOUND A =\_\_\_\_\_\_\_\_\_\_\_\_

**Mini-quiz #15 Chemistry 1114 Friday 11 October 2013**

**6 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ B**

1. Circle the compound formulas below that are not empiric formulas:

**C2 H2 O4 C2HO N21F7O3 H4S2O16**

1. Compound **B** contains31.17 % C, 62.33 % O and 6.493 % H by weight. Given that the atomic masses of C=12 g/mol, O=16 g/mol and H= 1 g/mol, determine the empiric formula for compound **B**.

EMPIRIC FORMULA OF COMPOUND B=\_\_\_\_\_\_\_\_\_\_\_\_

1. The molecular weight of compound **B** above is 385 g/mol. What is the molecular formula for compound **B**?

MOLECULAR FORMULA OF COMPOUND B =\_\_\_\_\_\_\_\_\_\_\_\_