**Mini-quiz #8 Chemistry 1114 Monday 16 September 2013**

Your name:\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ A

8.1. Write the complete, unabbreviated electronic configuration for:

1. S 1s22s22p63s23p4
2. K 1s22s22p63s23p64s1

8.2. Write the abbreviated electronic configurations for:

1. Be [He]2s2
2. Cl [Ne]3s23p5

8.3. Write the correct pigeonhole picture with d-switching and the filled, half-filled

empty rules applied to the element below (remember that +x means the

element has lost x electrons.)

1. Fe3+
2. Cu+

3d 4s

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Your name:\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ B

8.1. Write the complete, unabbreviated electronic configuration for:

1. P 1s22s22p63s23p3
2. Mg 1s22s22p63s2

8.2. Write the abbreviated electronic configurations for:

1. N [He]2s22p3
2. S [Ne]3s23p4

8.3. Write the correct pigeonhole picture with d-switching and the filled, half-filled

empty rules applied to the element below (remember that +x means the

element has lost x electrons.)

1. Fe+2
2. Cu

3d 4s