Mini-quiz 22 Chem 1013 Friday 26 April 2013 8 pts

Your name: \_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Balance us ! (2 points )

\_\_2\_\_C4H10 + \_\_13\_\_O2 🡪 \_\_8\_CO2 + \_10\_\_H2O

Given the balanced equation:

2Al + 6HCl 🡪 2AlCl3 + 3H2

How many grams of H2 form when 36 grams of Aluminum are consumed in the reaction ?

The molecular weight of H2 = 2 g/mol. The atomic weight of Aluminum is 27 g/mol. Show work !

Mol H2 = 3 = X (mol H2)

Mol Al 2 (36/27) mol Al

X= 2 mol H2= > weight H2 = mol H2 \* Mol. Wt H2 =2 \* 2 = 4

\_\_4\_\_\_\_\_\_ grams H2 formed