Mini-quiz 22 Chem 1013 Friday 26 April 2013 8 pts

Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Balance us ! (2 points )

\_\_\_\_C4H10 + \_\_\_\_O2 🡪 \_\_\_CO2 + \_\_\_H2O

Given the balanced equation:

2Al + 6HCl 🡪 2AlCl3 + 3H2

How many grams of H2 form when 36 grams of Aluminum are consumed in the reaction ?

The molecular weight of H2 = 2 g/mol. The atomic weight of Aluminum is 27 g/mol. Show work !

 \_\_\_\_\_\_\_\_ grams H2 formed

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