Mini-quiz 14 Chem 1013 11 March 2013 5 pts

Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

What is the pH of a solution with [H+]= 3.8\*10-4 M pH= \_\_\_\_\_-log[3.8\*10-4]= 3.42\_\_\_\_

What is the pOH of a solution with a pH of 5.6 ? pOH= \_14-5.6=8.4\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

What is [H+] if the pH is 5.6 ? [H+]=\_\_\_\_\_10-5.6 =2.51\*10-6\_\_\_\_\_\_\_M

If the pH is 1.5 what is the [OH-] of the solution [OH-]=\_\_\_\_\_3.16\*10-13\_\_\_\_\_\_\_\_\_\_\_\_M

pOH=14-1.5=12.5 =>[OH-]=10-12.5 = 3.16\*10-13

If the pOH is 3, is the solution acidic or basic ? BASIC ACIDIC (circle your choice)

pOH=3=> pH=14-3=11 basic