Chem 1013 mini-quiz 2 Monday 4 Sept 2017 A

Your name\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. List the three subatomic particles found in atoms (must get all right for credit)

\_\_electrons\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_protons\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_neutrons\_\_\_\_\_\_\_\_\_\_\_\_\_

2. Which is the lightest subatomic particle ?­­­­­­­­­­­­­\_\_\_\_\_electrons\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

3. Use your Periodic Table and fill out the table below: (2 pts)

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Atomic # | Mass # | symbol | #p | #n | #e | Atom charge |
| **14** | 30 | Si | 14 | 16 | 14 | 0 |

4. Write the atomic symbol ZMX for an isotope of Br (bromine) with 46 neutrons:

3581Br

5. Individual elements are distinguished by their count of \_\_\_\_\_protons\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 1013 mini-quiz 2 Monday 4 Sept 2017 B

Your name\_\_\_\_\_\_\_\_\_\_answers\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Which two subatomic particles have essentially the same mass ? \_\_\_\_protons\_\_\_\_\_ & \_\_neutrons\_\_\_\_\_\_\_\_

2. Use your Periodic Table and fill out the table below: (2 pts)

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Atomic # | Mass # | symbol | #p | #n | #e | Atom charge |
| **11** | 24 | Na | 11 | 13 | 11 | 0 |

3. Write the atomic symbol ZMX for an isotope of Si (silicon) with 15 neutrons: 1429Si

4.. Individual elements are distinguished by their count of \_\_\_\_\_\_protons\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

5. Mass number is defined as: \_\_\_\_\_proton count+ neutron count\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 1013 mini-quiz 2 Monday 4 Sept 2017 C

Your name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Use your Periodic Table and fill out the table below: (2 pts)

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Atomic # | Mass # | symbol | #p | #n | #e | Atom charge |
| **12** | 28 | Mg | 12 | 16 | 12 | 0 |

2. The mass of a neutron is:  **much greater** much smaller the same as the mass of an electron

(circle your choice)

3. Write the atomic symbol ZMX for an isotope of Fe (iron) with 31 neutrons: 2657Fe

4. Atomic number is defined as: \_\_\_\_\_\_proton count\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

5. Elements are defined by their \_\_\_\_\_\_proton count\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_