**Chem 1013: mini-quiz # 24: Redox concepts A 10 pts**

**Your name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

1a) oxidation means \_\_\_losing\_\_\_\_\_\_\_\_\_\_\_\_\_electrons

1b) reduction means\_\_\_gaining\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons

2) What is the oxidation number for each element (per element) in the compounds below:

H2O CO2 H2SO4

H=\_1\_\_ C=\_4\_\_ H=\_\_1\_ O=\_\_-2\_\_

O=\_-2\_\_ O=\_\_-2\_ S= \_6\_\_

Who is oxidized in the reaction below ?

Mgo + 2H+ 🡪 Mg2+ + H2

**Chem 1013: mini-quiz # 24: Redox concepts B 10 pts**

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1a) losing electrons is termed \_\_\_\_\_oxidation\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1b) gaining electrons is termed \_\_\_\_reduction\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2) What is the oxidation number for each element (per element) in the compounds below:

O2 H2O H2SO4 Cu2+

O=\_0\_ H=\_1\_\_ H=\_\_1\_ O=\_-2\_\_\_ Cu= \_\_2\_\_\_

O=\_-2\_\_ S= \_6\_\_

Who is oxidized in the reaction below ?

Zno + 2H+ 🡪 Zn2+ + H2