**Chem 1013: mini-quiz # 12: mole-mass-count word problem A 4 pts March 4**

Your name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Show work !

A lethal dose of the poison sodium cyanide is about 0.98 g for a 160 pound male named Bob. Given that the molecular weight of sodium cyanide is 49 g/mol, how many molecules of sodium cyanide does it take to kill Bob? Note that 1 mol count= 6.0\*1023 molecules.

 \_\_\_\_\_\_\_\_\_\_\_\_\_molecules sodium cyanide kills Bob

**Chem 1013: mini-quiz # 12: mole-mass-count word problem B 4 pts March 4**

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Show work !

Smoking two good medical marijuana joints is often prescribed as a way to reduce the nausea connected with cancer chemotherapy. The relevant chemical in the joints is THC (tetrahydrocannabinol). Given that the two joints deliver about 0.0629 g of THC , how many molecules of THC are being consumed? The molecular weight of THC=314.5 g/mol. 1 mol count =6\*1023 molecules

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_molecules of THC in 2 joints