**Homework #4: Chemistry 1013 Spring 2012**

**Due Wednesday 22 February 20 pts**

**4.1. Draw the best Lewis structure for H2SO3 (sulfurous acid) [ assume the S is in center of molecule surrounded by an O and two O-H (2 pts)**

**4.2. Determine the formal charges on each atom in the compounds drawn below:**

**(1 pt per charge/ 8 pts total )**

1. **:C**≡**O: formal charge on C and O ?**

**.. .. ..**

1. **:O-S=O: formal charge on O1, S and O2 ?**

**..**

**(O1) (O2)**

**.. .. ..**

1. **: O=N-O: formal charge on O1, N and O2 ? (note: this molecule has a net**

**.. charge of -1)**

**(O1) (O2)**

**4.3. Draw a reasonable organic compound that follows the HONC rules for the compound formulas below (more than one may be possible) 2 pts each / 8 pts total**

1. **C2H6O b) C2H4 c) C3H6O d) C2H5NO**

**4.4. Circle the atoms in the (bogus) structure below that fail to obey the HONC rules. (2 pts)**

H C O N-H

|

H

Cl